

1. CuCl_2
2. Al_2O_3
3. NiBr_3
4. Li_2S
5. B_2O_3
6. Pb_3N_2
7. BeF_2
8. BaCl_2
9. MnO_2
10. Fe_2S_3
11. AgBr
12. Nb_2O_5
13. Pt_3N_4
14. Cd_3P_2
15. CrSe_3

-These compounds do not require Roman Numerals, because their cations can only have one charge, making it unnecessary to indicate which charge to use.

-All of these compounds contain ionic bonds.

-The LCM represents the lowest value at which the positive and negative charges balance, creating a neutral compound.