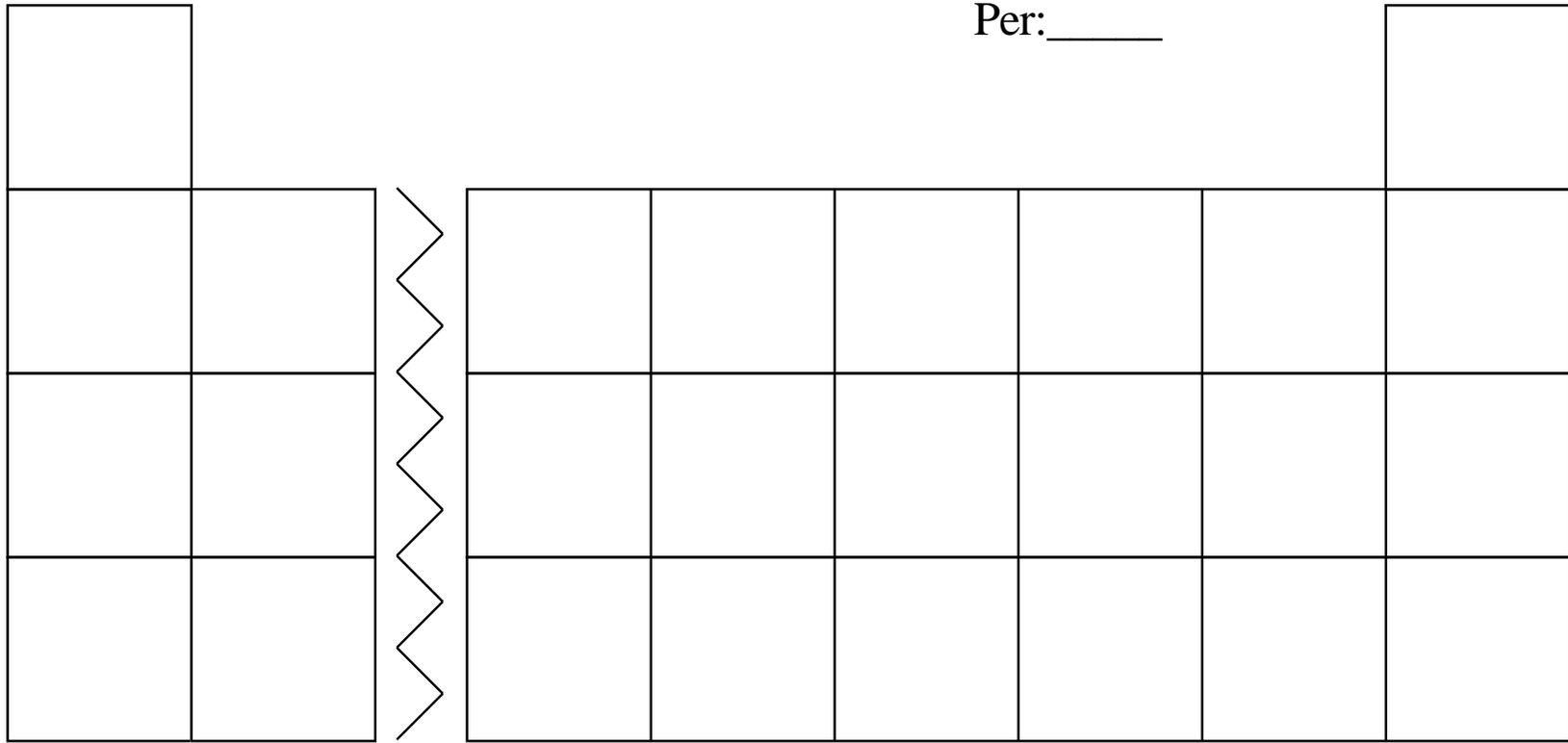


Name: _____

Per: _____



Conclusions:

-Discuss the general trends on the periodic table relating to atomic radii and electronegativity?

-Why are smaller atoms “stronger?”

-Metals are on the left side of the periodic table. As such, what do we know about their size, strength and number of valences electrons? As a result, how will metals behave when they react?

Instructions:

- *Cut out all the lettered boxes A-Z below
- *Use the following clues to place these boxes on the abbreviated periodic table on page 1
- *Glue the boxes in place and answer the conclusion questions

Clues:

- *The following elements are in groups: ZRD, PSIF, JXBE, LHT, QKA, WOV, GUN, YMC
- *J has an atomic number 3 times that of T
- *U has a total of six electrons
- *I₂A is the simple formula for an oxide
- *P is less dense than S
- *S in an alkali metal
- *E is a noble gas
- *W is a liquid
- *Z has the smallest atomic mass in its group
- *B has 10 protons
- *O has an atomic number larger than V
- *M has an atomic number one less than that of A
- *The atomic radius of K is the largest of the group
- *D has the largest atomic mass of its group
- *C has five electrons in its outer energy level
- *F is a gas
- *X has an atomic number one higher than F
- *L is an alkaline earth element with atomic mass of 40
- *Y is a metalloid
- *O is a halogen
- *The atomic mass of T is more than that of H
- *Q has an atomic mass 2 times that of A
- *Atoms of I are larger than those of S
- *The electrons of atom N are distributed over three energy levels

A	B	C	D	E	F	G	H	I
J	K	L	M	N	O	P	Q	R
S	T	U	V	W	X	Y	Z	